Term Test II Practice Test

1) Balance the following redox reaction if it occurs in acidic solution. What are the coefficients in front of $H_2C_2O_4$ and H_2O in the balanced reaction?

$$MnO_4^-(aq) + H_2C_2O_4(aq) \rightarrow Mn^{2+}(aq) + CO_2(g)$$

A) $H_2C_2O_4 = 5$, $H_2O = 8$
B) $H_2C_2O_4 = 1$, $H_2O = 1$
C) $H_2C_2O_4 = 5$, $H_2O = 1$
D) $H_2C_2O_4 = 1$, $H_2O = 4$
E) $H_2C_2O_4 = 3$, $H_2O = 2$
Answer: A

2) Balance the following redox reaction if it occurs in basic solution. What are the coefficients in front of ClO_2 and H_2O in the balanced reaction?

$$H_2O_2(l) + ClO_2(aq) \rightarrow ClO_2^-(aq) + O_2(g)$$